

Types of Reactions: Identifying and predicting products

For each reaction equation:

1. Balance the equation.
2. Identify the type of reaction (synthesis, decomposition, etc.).

Reaction Type

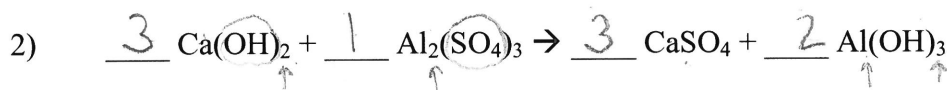
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|---|--------|
| 1. $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | 1. D |
| 2. $2 \text{KBr} + \text{F}_2 \rightarrow 2 \text{KF} + \text{Br}_2$ | 2. SR |
| 3. $1 \text{S}_8 + 24 \text{F}_2 \rightarrow 8 \text{SF}_6$ | 3. S |
| 4. $1 \text{Na}_2\text{SO}_4 + 1 \text{Pb}(\text{NO}_3)_2 \rightarrow 2 \text{NaNO}_3 + 1 \text{PbSO}_4$ | 4. DR |
| 5. $1 \text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$ | 5. C |
| 6. $1 \text{Zn} + 2 \text{HCl} \rightarrow 1 \text{ZnCl}_2 + 1 \text{H}_2$ | 6. SR |
| 7. $2 \text{FeCl}_3 + 3 \text{Na}_2\text{CO}_3 \rightarrow 1 \text{Fe}_2(\text{CO}_3)_3 + 6 \text{NaCl}$ | 7. DR |
| 8. $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + 1 \text{H}_2$ | 8. SR |
| 9. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ | 9. C |
| 10. $2 \text{HgO} \rightarrow 2 \text{Hg} + 1 \text{O}_2$ | 10. D |
| 11. $3 \text{CaCl}_2 + 2 \text{Na}_3\text{PO}_4 \rightarrow 6 \text{NaCl} + 1 \text{Ca}_3(\text{PO}_4)_2$ | 11. DR |
| 12. $1 \text{N}_2\text{O}_5 + 1 \text{H}_2\text{O} \rightarrow 2 \text{HNO}_3$ | 12. S |
| 13. $1 \text{Mg} + 1 \text{CuCl}_2 \rightarrow 1 \text{MgCl}_2 + 1 \text{Cu}$ | 13. SR |

Types of Reactions Worksheet

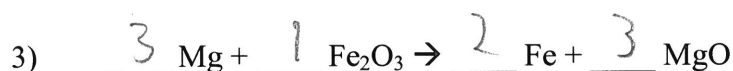
Balance the following equations and indicate the type of reaction taking place:



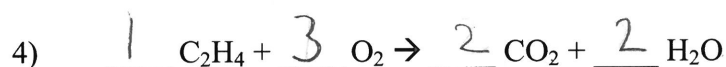
Type of reaction: DR



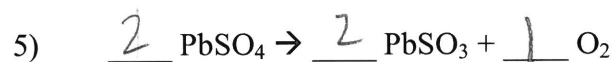
Type of reaction: DR



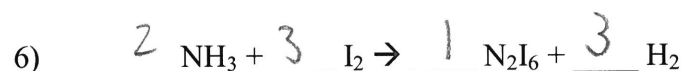
Type of reaction: SR



Type of reaction: C



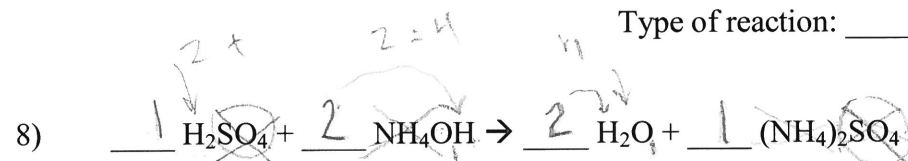
Type of reaction: D



Type of reaction: SR



Type of reaction: S



Type of reaction: N(DR)